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$4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ 4 a. What is the value of the coefficient x in this equation? 40.07 g/mol b. What is the molar mass of C_3H_4 ? 2 mol O_2 :1 mol H_2O c. What is the mole ratio of O_2 to H_2O in the above equation? 0.20 mol d. How many moles are in an 8.0 g sample of C_3H_4 ? 3z e. If z mol of C_3H_4 react, how many moles of ...

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Modern Chemistry Chapter 9 Mixed Review Answers CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield.

Answers Chapter 9 Mixed Review Modern Chemistry

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Modern Chemistry Chapter 11 GASES. Section 1. Gases & Pressure. Pressure & Force. Pressure (P) is defined as the force per unit of area on a surface. $\text{pressure} = \text{force} \div \text{area}$ $P = F/A$. Atmospheric pressure (atm) is the pressure exerted on an object due to the weight of the column of the air above it in the atmosphere. A .

Modern Chemistry Chapter 11 GASES

4? about 85°C b. What would be the boiling point of water if the air pressure over the liquid were reduced to 60 kPa? about 38 kPa c. What must the air pressure over CCl₄ be for it to boil at 50°C? d. Although water has a lower molar mass than CCl₄, it has a lower vapor pressure when measured? 20 STATES OF MATTER MODERN CHEMISTRY

10 States of Matter - Ms. Agostine's Chemistry Page

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