Chapter 9 Review Stoichiometry Section 3

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Chapter 9 Review Stoichiometry Section

CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N 2 are mixed with 12.0 mol of H 2 according to the following equation: N 2(g) 3H 2(g) →

2NH 3(g) N

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Modern Chemistry 77 Stoichiometry CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. _____ The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N 2 are mixed with 12.0 mol of H

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Stoichiometry. SECTION 1. SHORT ANSWER Answer the following questions in the space provided. 1. _____ The

coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products. (c) number of atoms of each element in each compound in a reaction.

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CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N 2 are mixed with 12.0 mol of H 2 according to the following

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percentage yield 2 60 mol of N 2 are mixed with 120 mol of H

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CHAPTER 9 REVIEW. Stoichiometry. SECTION 9.2. PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas: 2KClO3(s) (2KCl(s) + 3O2(g) How many grams of O2 form if 3.0 mol of KClO3 are totally consumed? 2. Given the following equation ...

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Reaction stoichiometry uses molar relationships to determine the amounts of unknown reactants or products from the amounts of known reactants or products. CHAPTER 9 DO NOT EDIT--Changes must be made through "File info" CorrectionKey=NL-A

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Review Module / Chapters 9–12 13 Prentice Hall, Inc. All rights In your notebook,solve the following problems. SECTION 9.1 THE ARITHMETIC OF EQUATIONS Use the 3-step problem-solving approach you learned in Chapter 4. 1. An apple pie needs 10 large apples,

2 crusts (top and bottom), and 1 tablespoon of cinnamon.

9 Stoichiometry Practice Problems

CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left Show all your work in the space provided 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g Calculate the percentage yield 2 60 mol of N 2 are mixed with 120 mol of H 2 according to the following equation: N 2(g) 3H 2(g

Download Chapter 9 Review Stoichiometry Section 2 Work

Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles of reactants/products involved in a chemical reaction. Students had an introduction to composition stoichiometry in Chapter 3 and will now move on to some more difficult problems.

Stoichiometry Worksheet Answers Chapter 9

CHAPTER 9 REVIEW Stoichiometry SECTION 2 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 4.5 mol The following equation represents a laboratory preparation for oxygen gas: 2KClO 3(s) → 2KCl(s) 3O 2(g) How many moles of O 2 form if 3.0 mol of KClO 3 are totally consumed? ...