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$\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)} + \text{Energy}$. Thermochemistry determine the heat exchanged at constant pressure, $q = m c \Delta T$. Calculating the limiting reactant, the change in enthalpy of the reaction, ΔH_{rxn} , can be determined since the reaction was conducted under conditions of constant pressure $\Delta H_{\text{rxn}} = q_{\text{rxn}} / \# \text{ moles of limiting reactant}$. This reaction is classified as an ...

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